

## Chapter 7 Ionic Metallic Bonding Work Answers

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### Chapter 7 Ionic Metallic Bonding

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Section 7.3 – Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

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### **Chapter 7 - Ionic and Metallic Bonding**

Chapter 7 - Ionic and Metallic Bonding. Jennie L. Borders. Section 7.1 - Ions. Valence electrons are the electrons in the highest occupied energy level. Valence electrons are the only electrons involved in chemical bonding. Elements in the same group have the same number of valence electrons.

### **Chapter 7 - Ionic and Metallic Bonding**

Chapter 7 "Ionic and Metallic Bonding ... Ionic Bonding. Anions and cations are held together by opposite charges (+ and -) Ionic compounds are called salts. Formula unit - Simplest ratio of elements in an ionic compound. Practice with Formula Unit. What is the formula unit for the following compounds?

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Chapter 7 Ionic/Metallic Bonding. the electrons involved in metallic bonding that are free to move easily from one atom to the next throughout the metal and are not attached to a particular atom. proposes that all metal atoms in a metallic solid contribute their valence electrons to form a "sea" of electrons, and can explain properties of metallic solids such as malleability, conduction, and ductility.

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Chapter 7: Ionic and Metallic Bonding. valence electrons. electron dot structures. octet rule. halide ions. electrons in the highest occupied energy level of an element's.... diagrams that show valence electrons in the atoms of an element.... in forming compounds, atoms tend to achieve the electron configuration....

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### **BONDING AND INTERACTIONS**

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